

Ch. 14- Chemical Kinetics

Topics for Notes	Video Title	Video Link	Length	Practice Problems	Pages
Definition of reaction rates, average rate versus instantaneous rate,	The Rate of Reactions	http://www.bozeman-science.com/ap-chem-035-the-rate-of-reactions	6:24	FP 14.1, CC 14.1	623-628
Rate law, reactant order, first order, second order, third order, overall order	The Rate Law	http://www.bozeman-science.com/ap-chem-036-the-rate-law	8:43	CC 14.2, CC 14.3	629-634 (stop at section 14.4)
Rate constant	The Rate Constant	http://www.bozeman-science.com/ap-chem-037-the-rate-constant	6:42	FP 14.2	629 (skip pages 635-641)
Draw diagram from figure 14.12 with activation energy labeled, activated complex	The Reactions Path	http://www.bozeman-science.com/ap-chem-040-the-reaction-path	5:36	CC 14.7	642-643 (skip 644-646)
Collision Model	Activation Energy	http://www.bozeman-science.com/ap-chem-039-unsuccessful-reactions	4:51	Draw the 3 pictures with labels from the bottom of page 647 and explain why each is effective or ineffective, CC 14.8	647-648

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Elementary step	Elementary Reactions	http://www.bozeman-science.com/ap-chem-038-elementary-reactions	6:12	#75	649
Reaction intermediate, unimolecular, bimolecular, termolecular	Intermediate Steps	http://www.bozeman-science.com/ap-chem-043-reaction-intermediates	3:28	#77	649-650
Rate-determining step	Rate-Determining Step	http://www.bozeman-science.com/ap-chem-042-the-rate-limiting-step	6:42	FP 14.9, SAQ 11	650-652
Catalyst	Catalysts	http://www.bozeman-science.com/ap-chem-044-catalysts	4:09		653-654
Homogeneous catalysis, heterogeneous catalysis	Catalyst Classes	http://www.bozeman-science.com/ap-chem-045-catalyst-classes	6:09	#79	655-658