Ch. 15- Chemical Equilibrium, LeChatelier's, & Ksp

Topics for Notes	Video Title	Video Link	Length	Problems	Pages
Dynamic equilibrium, reversible	Crash Course: Equilibrium	https://www.youtube.com/watch?v=g5wNg_dKsYY	10:39	25	675-679
1. Equilbrium constant (K- notice it's a capital K not a lowercase one like in kinetics, yay) 2. Write and label the parts of an equilibrium expression like at the bottom of page 680	Crash Course: Equilibrium Equations	https://www.youtube.com/watch?v=DP-vWN1yXrY	9:28	FP 15.1	680
1. Copy "Summarizing the Significance of the Equilibrium Constant" on p. 681 (I remember it by little K's lie to the left- as in they favor reactants) 2. Copy and describe 3 rules for equilibrium constant and chemical equation on p. 683	The Equilibrium Constant	http://www.bozemanscience.com/ap-chem-065-the-equilibrium-constant	6:14	CC 15.1, CC 15.2, CC 15.3, FP 15.2, FMP 15.2, SAQ #1	680-684 and 687-688 -Skip sections 15.4-15.8
LeChatelier's Principle Note: Q is calculated just like the equilibrium constant value BUT with	LeChatelier's Principle	http://www.bozemanscience.com/ap-che6-066-lechateliers-	7:00	FP 15.14, FP 15.15, FP 15.16, SAQ #11	703-710

molarities from before the reaction is at equilibrium		principle			
K_{sp}	Solubility	http://www.bozemanscience.com/ap-chem-070-solubility	7:03	Read p.811 Hard Water – Show work for the "Question- $CaCO_3$" CC 17.9	809-813 *hint: solubility is equal to x in the ice charts