Ch. 17- Aqueous Ionic Equilibrium (Great intro video to buffers!) https://www.youtube.com/watch?v=NJyAme5GVF8

Topics	Video Title	Video Link	Length	Problems	Pages
Buffer, common ion effect,	Common Ion Effect	https://www.youtube.com/watch?v=NPa6eN3iMZU	10:43	FP 17.1	779-783
Henderson-Hasselbach Equation (it's on the formula sheet),	Buffers: the Acid Rain Slayer- Crash Course #31	https://www.youtube.com/watch?v=8Fdt5WnYn1k	11:40	FP 17.2, CC 17.2, CC 17.3, 17.4	783-790
Qualities of good buffers	pH and Buffers	http://www.bozemanscience.com/ap-chem-069-ph	5:56	@2:47 in video write equilibrium expression for a weak acid @3:05+ in video write down 2 factors that make a good buffer	791-794
Titrations and pH curves	Acid-Base Equilbrium	http://www.bozemanscience.com/ap-chem-068-equilibrium-reasoning	Watch starting at 6:30 *until end* (about 4 minutes)	1. Draw the three titration curves (graphs that look like an "s") for a strong acid/strong base, strong B/weak A, strong A/weak B*, and label them. *pH is lower than 7 here because conjugate acid of weak base is formed during titration 2. BESIDE EACH write one sentence describing what is going on with the data in the curve. No reading from your text needed, but it is section 15.4 if you want to look.	