

# I. Metric Mania Lab

## II. Theory and Purpose:

Concepts studied: metrics, scientific method, sig figs, DA

How do you measure with scientific instruments?

**Instruments:**                      **Accurate to:**

Graduated cylinder	0.0 mL
Digital Balance	0.0 g
Metric ruler	0.0 cm

III. Prelab: (use this for help: <http://www.molelady.com/html/METRICHO.htm> )

1. What is the basic unit for length? \_\_\_\_\_
2. What is the basic unit for mass? \_\_\_\_\_
3. What is the basic unit for temperature? \_\_\_\_\_
4. What are the freezing and boiling points for water on this scale? \_\_\_\_\_
5. What is the basic unit for volume? \_\_\_\_\_
6. What is the best unit for measuring each distance:
  - a. Thickness of an eyelash:    mm    cm    m
  - b. Length of a pencil:    cm    m    km
7. What is the best unit for measuring each mass:
  - a. Amount of spices in a batch of cookies:    mg    g    kg
  - b. Your mass:    mg    g    kg
  - c. Mass of 10 pennies:    mg    g    kg
8. What is the best unit for measuring each volume?

Amount of soda in 1 can:    mL            L

Water in a bathtub:    mL            L
9. Circle the best choice: (all temperatures are in Celsius)
  - a. Temperature on a hot summer's day: 0°    35°    90°
  - b. Room temperature: - 20°    0°    20°

## Procedure:

Explore measurements of length, mass, volume, and temperature. Use the accuracy of the instrument to record a unit of measurement for each answer. Remember to record the APPROPRIATE amount of sig figs.

**IV. Data:** Use units of measurement and sig figs in your lab notebook with appropriate headings.

### **Part A. LENGTH:**

In the lab #   1  , measure the circumference of a balloon.(cm) \_\_\_\_\_ cm.    Then convert to km.

Show your work here:

In the lab #   7  , use a metric ruler to find each measurement.

- a. Width of this page \_\_\_\_\_ mm or \_\_\_\_\_ cm
- b. Length of an unsharpened pencil \_\_\_\_\_ cm

**Part B. MASS: Remember to include the accuracy of the instrument as part of the sig figs in your answer.**

In lab #   2  , use a digital balance to find these masses.

- a. Mass of an ink pen \_\_\_\_\_ g      b. Mass of a can of soda \_\_\_\_\_ g

**Part C. TEMPERATURE:**

At Lab #   3  

Record time you took the temp: \_\_\_\_\_

Take the temperature of the water. \_\_\_\_\_

Convert to K. \_\_\_\_\_ by using  $K = ^\circ C + 273$

\*\*\*After you take the temperature, put the thermometer or probe back where you found it.

**Part D. VOLUME:**

At Lab #   4  ,

- a. Use the eraser and record the following: (don't forget the units of measurement)

Length: \_\_\_\_\_

Width: \_\_\_\_\_

Height: \_\_\_\_\_

Use  $L \times W \times H$  to find the volume of a chalkboard eraser \_\_\_\_\_  $\text{cm}^3$

- b. Convert to mL \_\_\_\_\_

At Lab #   5  , Identify the name of each glassware and read the volume:

A. Item 1 \_\_\_\_\_ Volume \_\_\_\_\_

B. Item 2 \_\_\_\_\_ Volume \_\_\_\_\_

C. Item 3 \_\_\_\_\_ Volume \_\_\_\_\_

At Lab #   6  , Measure 10.0 mL in the following pieces of glassware & record the measurement.

A. Graduated cylinder (accurate to .0 mL) \_\_\_\_\_ mL

B. A small beaker (accurate to the whole number) \_\_\_\_\_ mL

C. Test tube (1 mL = 10 drops) \_\_\_\_\_ mL

\*\*\*\* Put your water sample into the sink. Put glassware back into drawers 1 or 2.

V. PostLab

Calculations: Show your work below each measurement – Do this in PENCIL!

1. Convert the following measurements. (use  $K = ^\circ C + 273$ )

a.  $90.5^\circ C = \underline{\hspace{2cm}} K$

b.  $45 K = \underline{\hspace{2cm}} ^\circ C$

2. Convert the following measurements: **(by DA - show all work)**

a.  $1.6 \times 10^{-3} \text{ mg} = \underline{\hspace{2cm}} \text{ g}$

b.  $4.7 \text{ kg} = \underline{\hspace{2cm}} \text{ g}$

3. Convert the following measurements: **(by DA - show your work for each)**

a.  $34 \text{ mm} = \underline{\hspace{2cm}} \text{ cm}$

b.  $3 \text{ km} = \underline{\hspace{2cm}} \text{ m}$

4. Convert the following measurements: **(by DA)**

a.  $2.30 \times 10^4 \text{ KL} = \underline{\hspace{2cm}} \text{ L}$

b.  $456 \text{ cL} = \underline{\hspace{2cm}} \text{ mL}$

c.  $20. \text{ mL} = \underline{\hspace{2cm}} \text{ cm}^3$