

*Naming and Writing Chemical Formulas Review*  
*AP Chem*

**Write the correct formula for the compound formed by each of the following pairs of ions.**

- |  |           |
|--|-----------|
| 1. Na <sup>1+</sup> F <sup>1-</sup>                | 1. _____  |
| 2. K <sup>1+</sup> S <sup>2-</sup>                 | 2. _____  |
| 3. Ni <sup>2+</sup> SO <sub>4</sub> <sup>2-</sup>  | 3. _____  |
| 4. Al <sup>3+</sup> O <sup>2-</sup>                | 4. _____  |
| 5. Ca <sup>2+</sup> ClO <sub>3</sub> <sup>1-</sup> | 5. _____  |
| 6. NH <sub>4</sub> <sup>1+</sup> P <sup>3-</sup>   | 6. _____  |
| 7. Cu <sup>1+</sup> NO <sub>3</sub> <sup>1-</sup>  | 7. _____  |
| 8. Cu <sup>2+</sup> NO <sub>3</sub> <sup>1-</sup>  | 8. _____  |
| 9. Pb <sup>4+</sup> O <sup>2-</sup>                | 9. _____  |
| 10. Li <sup>1+</sup> CO <sub>3</sub> <sup>2-</sup> | 10. _____ |

**For each of the following compounds, write A) the symbols of the ions in the compound and the B) chemical name. Do you need the list of monatomic and polyatomic ions? It is [here](#).**

	Symbols of ions	NAME
11. CaI <sub>2</sub>	11. _____	_____
12. Na <sub>2</sub> CO <sub>3</sub>	12. _____	_____
13. Ga(ClO <sub>3</sub> ) <sub>3</sub>	13. _____	_____
14. CuF <sub>2</sub>	14. _____	_____
15. (NH <sub>4</sub> ) <sub>3</sub> PO <sub>4</sub>	15. _____	_____
16. FeSO <sub>4</sub>	16. _____	_____
17. Mg(NO <sub>3</sub> ) <sub>2</sub>	17. _____	_____
18. NH <sub>4</sub> NO <sub>2</sub>	18. _____	_____
19. KC <sub>2</sub> H <sub>3</sub> O <sub>2</sub>	19. _____	_____
20. Na <sub>2</sub> Cr <sub>2</sub> O <sub>7</sub>	20. _____	_____